**2.1 Surface O3 chemical budget (160)**

The surface O3 chemical budget involves the consideration of production and loss, which are represented by P(O3) and L(O3), respectively, while excluding the deposition and atmospheric dynamic-driven fluxes. The production rate of O3 is equivalent to the instantaneous photolysis of NO2 when subjected to radiation of wavelength <330 nm. The production of NO2, on the other hand, originates from the oxidation of NO by HO2, CH3O2, and RO2 peroxide radicals. This results in the O3 production formula, as described in Equation 1. The destruction of O3 occurs through photolysis under radiation of wavelength <330 nm, reactions with HOX radicals, and unsaturated VOCs, such as alkenes, as indicated in Equation 2. The temperature-dependent kinetic rate coefficients are determined based on the atmospheric chemistry module's configuration coupled into the Earth system model, for example, UKCA1 embedded in UKESM1-0-LL2. These values are calculated using the simulated surface air temperature and are detailed in Supplementary Table 1, with reference values evaluated by IUPAC.

NO2 + *hv* → NO + O(3P) (R1)

O(3P) + O2 + M → O3 + M (R2)

NO + HO2 → NO2 + OH (R3)

NO + CH3O2 → NO2 + CH3O (R4)

NO + RO2 → NO2 + RO (R5)

O3 → O2 + O(1D) (R6)

O(1D) + H2O → 2OH (R7)

O3 + OH → O2 + HO2 (R8)

O3 + HO2 → 2O2 + OH (R9)

P(O3) = *k*1[NO][HO2] + *k*2[NO][CH3O2] + ∑*ki*[NO][RO2] (Equation 1)

L(O3) = *k*3[O(1D)][H2O] + *k*4[O3][OH] + *k*5[O3][HO2] + ∑*kj*[O3][VOCs] (Equation 2)

, preferred value as 8.5×10–12 cm3 molecule-1 s-1 at 298 K.

, preferred value as 7.7×10–12 cm3 molecule-1 s-1 at 298 K.

, preferred value as 2.14×10–10 cm3 molecule-1 s-1 independent of temperature over the range 200–350 K.

, preferred value as 7.3×10–14 cm3 molecule-1 s-1 at 298 K.

, preferred value as 2.0×10–15 cm3 molecule-1 s-1 at 298 K.

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